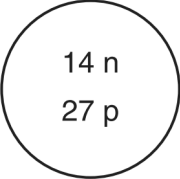
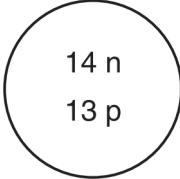
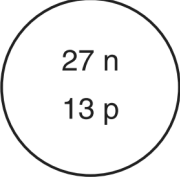
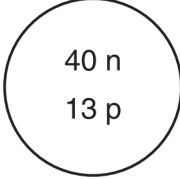


Ch 3: Level 1 Questions

1. An atom of $^{226}_{88}\text{Rn}$ contains
- A. 88 protons and 138 neutrons
 - B. 88 protons and 138 electrons
 - C. 88 electrons and 226 neutrons
 - D. 88 electrons and 226 protons
2. Experimental evidence indicates that the nucleus of an atom
- A. contains most of the mass of the atom
 - B. contains a small percentage of the mass of the atom
 - C. has no charge
 - D. has a negative charge
3. An atom that contains 8 protons, 8 electrons, and 9 neutrons has
- A. an atomic number of 9
 - B. an atomic number of 16
 - C. a mass number of 17
 - D. a mass number of 25
4. Compared to the entire atom, the nucleus of the atom is
- A. smaller and contains most of the atom's mass
 - B. smaller and contains little of the atom's mass
 - C. larger and contains most of the atom's mass
 - D. larger and contains little of the atom's mass
5. What is the total number of protons and neutrons in an atom of $^{86}_{37}\text{Rb}$?
- A. 37
 - B. 49
 - C. 86
 - D. 123
6. An atom of carbon-12 and an atom of carbon-14 differ in
- A. atomic number
 - B. mass number
 - C. nuclear charge
 - D. number of electrons
7. The nucleus of an atom of cobalt-58 contains
- A. 27 protons and 31 neutrons
 - B. 27 protons and 32 neutrons
 - C. 59 protons and 60 neutrons
 - D. 60 protons and 60 neutrons
8. Which diagram represents the nucleus of an atom of $^{27}_{13}\text{Al}$?
- A. 
- B. 
- C. 
- D. 
9. Which isotopic notation represents an atom of carbon-14?
- A. ^6_8C
 - B. ^8_6C
 - C. $^{14}_6\text{C}$
 - D. $^{14}_8\text{C}$
10. The notation for the nuclide $^{137}_{55}\text{Cs}$ gives information about
- A. mass number, only
 - B. atomic number, only
 - C. both mass number and atomic number
 - D. neither mass number nor atomic number

1.
Answer: A
2.
Answer: A
3.
Answer: C
4.
Answer: A
5.
Answer: C
6.
Answer: B
7.
Answer: A
8.
Answer: B
9.
Answer: D
10.
Answer: C