

Unit 5 Test Review

(Gen)

Name: _____

**Note: There are multiple choice questions on the Unit 5 Test. You can practice for these using the two posted Quia activities.

Write the formula:

- 1) Iron (III) sulfate
- 2) Lithium nitride
- 3) Disulfur decafluoride
- 4) Platinum IV sulfide
- 5) Nickel III phosphate
- 6) Dinitrogen tetroxide
- 7) Copper (II) phosphide
- 8) Tetraphosphorus heptoxide
- 9) Magnesium carbonate
- 10) Aluminum hydroxide
- 11) Dichlorine octoxide

$\text{Fe}_2(\text{SO}_4)_3$
 Li_3N
 S_2F_{10}
 PtS_2
 NiPO_4
 N_2O_4
 Cu_3P_2
 P_4O_7
 MgCO_3
 $\text{Al}(\text{OH})_3$
 Cl_2O_8

Name the compound:

- 12) CCl_4
- 13) NH_4OH
- 14) $\text{Co}_2(\text{SO}_4)_3$
- 15) P_2O_5
- 16) Mn_2S_3
- 17) KBr
- 18) Cu_2CO_3
- 19) PCl_3
- 20) $\text{Fe}_3(\text{PO}_4)_2$
- 21) N_2S_3
- 22) $\text{Al}(\text{NO}_3)_3$
- 23) Ag_3N

Carbon tetrachloride
ammonium hydroxide
cobalt III sulfate
diphosphorus pentoxide
manganese III sulfide
potassium bromide
copper I carbonate
phosphorus trichloride
iron II phosphate
dinitrogen trisulfide
aluminum nitrate
silver I nitride

Short Answer:

- 24) What happens to electrons in an ionic bond? transferred
- 25) What happens to electrons in a covalent bond? shared
- 26) What two components make up an ionic bond? cation + anion, or metal + nonmetal
- 27) What two components make up a covalent bond? nonmetal + nonmetal
- 28) Why do atoms form bonds? to be stable, 8 valence electrons (octet)
- 29) What is the strongest type of bond? ionic
- 30) What holds the atoms of an ionic bond together? magnetic attraction
- 31) When do you use roman numerals? transition metals
- 32) Is the compound NH_4OH ionic or covalent? Why? ionic: cation NH_4^+ anion: OH^-
- 33) What type of bond would an element from Group 1 be likely to form with an element from Group 17? ionic
- 34) What type of bond would an element from Group 16 be likely to form with an element from Group 17? covalent