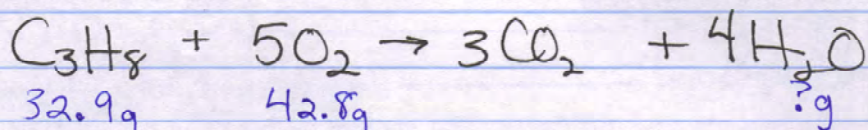


Limiting Reactant Problems

Step 1: write a balanced equation



Step 2: write the amounts given, for each reactant under the appropriate compound

Step 3: Choose one product to solve for

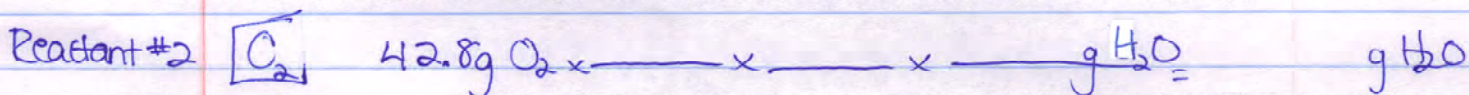
Note: it does not matter which product you choose, unless the problem specifies (most problems do not)

I will pick H_2O to solve for.

I will put a question mark underneath it in the problem to remind me.

Step 4: Solve for grams of the product you chose. Start with a different reactant given in the problem.

(AKA - you will have two gram-to-gram conversions)



Step 5: The reactant that produces the smallest amount of the product chosen, is the limiting reactant. The other reactant is the excess reactant.

The smallest amount of product_n is the amount of product that will actually be made because the limiting reactant will run out first & "limit" the amount of product made. You would have excess reactant left over, but it doesn't matter because there is no more limiting reactant left!

