

**Comprehension Questions for Intermolecular Forces Video**

Name: \_\_\_\_\_

- 1) Intermolecular forces are the forces \_\_\_\_\_.
- 2) The attractive forces between molecules is what governs all \_\_\_\_\_ properties.
- 3) Name two examples of physical properties:  
\_\_\_\_\_ and \_\_\_\_\_
- 4) Physical properties are governed by how strongly the molecules are \_\_\_\_\_.
- 5) What are the 3 types of intermolecular forces?  
\_\_\_\_\_, \_\_\_\_\_, \_\_\_\_\_
- 6) The degree of intermolecular forces is determined by how \_\_\_\_\_ or how \_\_\_\_\_ the charges on molecules are.
- 7) The symbol  $\delta^+$  stands for \_\_\_\_\_ and the symbol  $\delta^-$  stands for \_\_\_\_\_.
- 8) Polar molecules have a permanent \_\_\_\_\_. (which means a permanent partial positive or partial negative charge).
- 9) What type of interaction is stronger than a dipole-dipole interaction? \_\_\_\_\_
- 10) Hydrogen bonding occurs when Hydrogen is bonded to a highly \_\_\_\_\_ element. These three elements are \_\_\_\_\_, \_\_\_\_\_, and \_\_\_\_\_.
- 11) Non-polar molecules have no \_\_\_\_\_. (meaning no permanent partial positive or negative charge)
- 12) What type of intermolecular force is observed when a nonpolar carbon tetrachloride molecule comes in contact with another carbon tetrachloride molecule? \_\_\_\_\_
- 13) Dispersion forces are based on a \_\_\_\_\_ dipole moment.
- 14) Which is stronger? Circle your answer: dipole-dipole interaction or dispersion forces

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