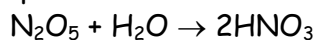


Chapter 11: Mole-to-Mole Conversions & Review Name: _____

Stoichiometry

Mole-to-Mole Conversions

- 1) How many moles of HNO₃ will be produced when 0.51 mol of N₂O₅ reacts?



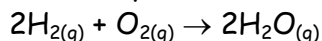
1.02 mol

- 2) How many moles of NaBr will be produced when 0.69 mol of bromine (Br₂) reacts according to the following equation: $\text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + \text{I}_2$

1.38 mol

Mixture of Mass and Volume Conversions

- 3) How many liters of oxygen would be required to react completely with 250.0 L of hydrogen gas?



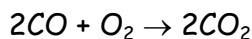
125 L

- 4) If 2.50 g silver reacts in the following equation, how many grams of HNO₃ will be used up?



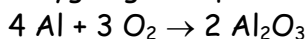
2.92 g

- 5) What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide? (Write a balanced equation first!)



0.44 g

- 6) Calculate the volume in liters of oxygen gas required to react with 50.0 g of aluminum at STP.



31.13 L