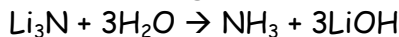


Chapter 11: Gram-to-Gram Conversions Name: _____

Stoichiometry

- 1) What mass of LiOH is produced when 0.38g of lithium nitride reacts?



0.78g

- 2) What mass of hydrogen peroxide must decompose according to the equation below to produce 0.77g of water?



1.45g

- 3) For the reaction $4\text{FeS} + 7\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{SO}_2$ how many grams of Fe_2O_3 can be formed if we start with 50.3 g of FeS?

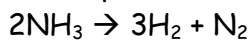
45.69g

- 4) If 2.50 g silver reacts in the following equation, how many grams of HNO_3 will be used up?



2.92g

- 5) What mass of ammonia must decompose to produce 0.87 g of hydrogen?

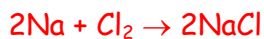


4.89g

- 6) What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium?

Step 1: Write a balanced equation

Step 2: Solve the gram to gram conversion



0.74g