Ch 7: Distinguishing between Ionic and Covalent Compounds (PrelB)

Name: _____

Write the **formula** of the following compounds:

in box: Ionic(I) Covalent(C)

I	1)	Sodium iodide	NaI
C	2)	Nitrogen trifluoride	NF_3
I	3)	Strontium fluoride	SrF_2
I	4)	Lead (II) sulfide	PbS
C	5)	Silicon tetrafluoride	SiF_4
C	6)	Iodine pentafluoride	IF_5
I	7)	Silver (I) carbonate	Ag_2CO_3
I	8)	Magnesium sulfate	$MgSO_4$
I	9)	Ammonium chloride	NH_4Cl
I	10)	Beryllium hydroxide	$Be(OH)_2$
C	11)	Sulfur tetrafluoride	SF_4

Write the **name** of the following compounds:

I	12)	K_2SO_3	Potassium sulfite
I	13)	Rb_2S	Rubidium sulfide
I	14)	CaF_2	Calcium fluoride
C	15)	SF_6	Sulfur hexafluoride
I	16)	ZnO	Zinc II oxide *(focus on the charge of oxygen in this empirical formula)*
I	17)	AgCl	Silver I chloride
I	18)	$NaNO_2$	Sodium nitrite
I	19)	Cu_2CO_3	Copper I carbonate *(focus on who the subscript of 3 "belongs to")*
I	20)	AlF_3	Aluminum fluoride
C	21)	NO	Nitrogen monoxide
I	22)	Ti_3P_2	Titanium II phosphide
C	23)	PCl_5	Phosphorus pentachloride
I	24)	$(NH_4)_3PO_4$	Ammonium phosphate
I	25)	$NiSO_4$	Nickel II sulfate *(focus on the charge of sulfate in this empirical formula)*