

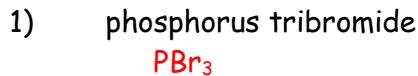
Ch 7: Covalent Compounds

Name: _____

Molecular (AKA "Covalent") Compounds

Indicate formula:

Indicate name:



Short Answer

13) Explain the difference between ionic and covalent bonds.

Ionic bonds = electrons transferred, magnetic attraction between cation & anion, strongest bond
Covalent bonds = electrons shared, between nonmetal & nonmetal

14) What are valence electrons?

the outermost electrons (or electrons in the highest energy level)

15) From where to atoms gain/ lose electrons?

valence electrons (or the highest energy level)

16) List the ten prefixes for molecular compounds. 1= **mono** 2= **di** 3= **tri**
4= **tetra** 5= **penta** 6= **hexa** 7= **hepta** 8= **octa** 9= **nona** 10= **deca**

Mixed Practice: Identify each compound as ionic (I) or covalent (C) = AKA molecular compounds

Write I or C in the box. You do not need to do anything else.

17) I lithium sulfide

22) C NO_2

18) C diphosphorus pentoxide

23) I AlF_3

19) I magnesium nitrate

24) I KNO_3

20) I nickel (III) phosphide

25) I Ga_2O_3

21) C silicon dioxide

26) C B_2O_3

