Ionic Compounds with Transition Metals & Polyatomic Ions

Name: _____

Ionic Compounds with Transition Metals

Indicate formula:

1) cobalt (III) oxide
   Co_2O_3

2) iron (II) bromide
   $FeBr_2$

3) silver (I) nitride
   Ag_3N

4) zirconium (III) phosphide
   ZrP

5) lead (IV) sulfide
   PbS_2

6) chromium (III) nitride
   CrN

Indicate name:

7) $FeCl_2$
   Iron II chloride

8) Ni_2S_3
   Nickel III sulfide

9) CrF_2
   Chromium II fluoride

10) TiI_4
    Titanium IV iodide

11) Co_3N_2
    Cobalt II nitride

12) Sn_3P_2
    Tin II phosphide

Ionic Compounds with Polyatomic Ions

Indicate formula:

13) magnesium nitrate
    $Mg(NO_3)_2$

14) calcium phosphate
    $Ca_3(PO_4)_2$

15) ammonium sulfide
    $(NH_4)_2S$

16) silver (I) hydroxide
    $AgOH$

17) chromium (III) carbonate
    $Cr_2(CO_3)_3$

18) aluminum hydroxide
    $Al(OH)_3$

Indicate name:

19) Na_2CO_3
    Sodium carbonate

20) K_2SO_4
    Potassium sulfate

21) $Zn(OH)_2$
    Zinc II hydroxide

22) NH_4OH
    Ammonium hydroxide

23) Li_3PO_4
    Lithium phosphate

24) $Fe(NO_3)_3$
    Iron III nitrate