

Ch 7: Binary Ionic Compounds

Name: _____

Indicate formula:

1) sodium phosphide



2) aluminum sulfide



3) beryllium chloride



4) potassium oxide



5) barium fluoride



6) gallium nitride



7) calcium oxide



8) magnesium selenide



Indicate name:

9) MgBr_2

Magnesium bromide

10) KCl

Potassium chloride

11) Al_2O_3

Aluminum oxide

12) Ca_3P_2

Calcium phosphide

13) GaF_3

Gallium fluoride

14) Sr_3N_2

Strontium nitride

15) RbI

Rubidium iodide

16) Na_2O

Sodium oxide

17) Explain the difference between cations and anions.

Cations have a positive charge because they have lost electrons.

Anions have a negative charge because they have gained electrons.

18) State the octet rule.

Atoms are stable when they have 8 valence electrons.

19) What two sublevels are involved in using the octet rule? s and p sublevels

20) Indicate the number of valence electrons and charge for each element.

	S	Cl	Ar	Ca	Ne	N
valence e^-	6	7	8	2	8	5
Charge	-2	-1	0	+2	0	-3

